Friday Worksheet   Volumetric 4
Name: ……………….

0.326 g of a pure acid, $H_2X(s)$, reacts with exactly 100 mL of 0.105 M NaOH(aq).
A reaction occurs according to the equation
$H_2X(s) + 2NaOH(aq) \rightarrow Na_2X(aq) + 2H_2O(l)$
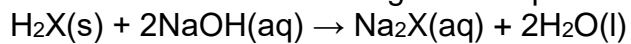
a. Calculate:
i. the amount, in mol, of NaOH that is added to the acid H_2X.

ii. the amount, in mol, of acid H_2X.

iii. the molar mass, in g mol⁻, of the acid H_2X

b. Identify acid H_2X