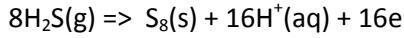
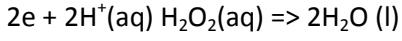
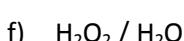
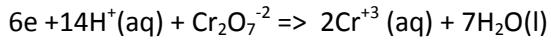
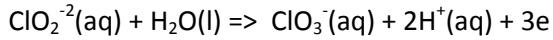
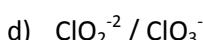
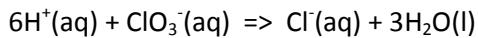
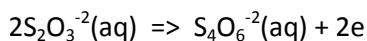
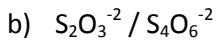
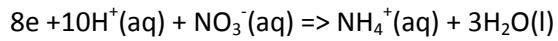
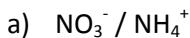
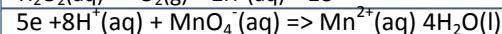
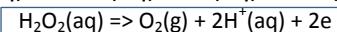
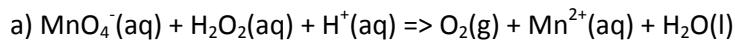


Revision/solutions

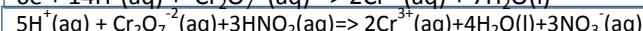
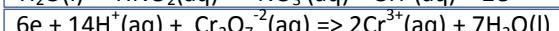
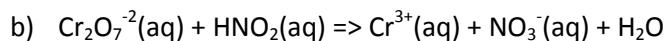
1) Write the balanced half equations to the following



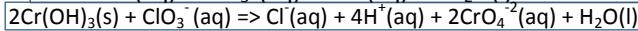
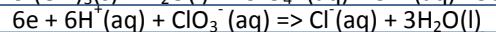
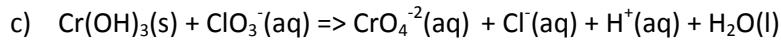
2) Below are the unbalanced redox reactions. Give the balanced oxidation and reduction half reactions for each and the balanced overall equation.



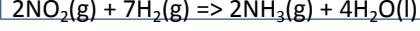
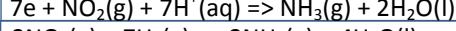
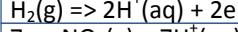
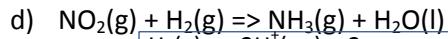
oxidation
reduction
overall



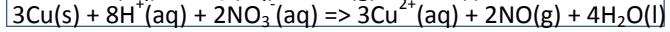
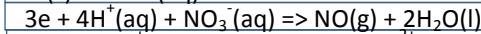
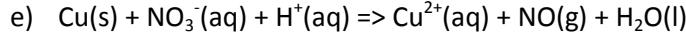
oxidation
reduction
overall



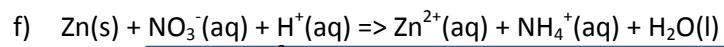
oxidation
reduction
overall



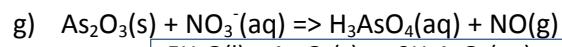
oxidation
reduction
overall



oxidation
reduction
overall



$Zn(s) \Rightarrow Zn^{2+}(aq) + 2e$	oxidation
$8e + 10H^+(aq) + NO_3^-(aq) \Rightarrow NH_4^+(aq) + 3H_2O(l)$	reduction
$4Zn(s) + 10H^+(aq) + NO_3^-(aq) \Rightarrow 4Zn^{2+}(aq) + NH_4^+(aq) + 3H_2O(l)$	overall



$5H_2O(l) + As_2O_3(s) \Rightarrow 2H_3AsO_4(aq) + 4H^+(aq) + 4e$	oxidation
$3e + 4H^+(aq) + NO_3^-(aq) \Rightarrow NO(g) + 2H_2O(l)$	reduction
$7H_2O(l) + 3As_2O_3(s) + 4H^+(aq) + 4NO_3^-(aq) \Rightarrow 4NO(g) + 6H_3AsO_4(aq)$	overall