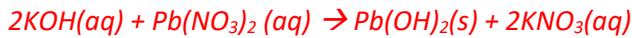


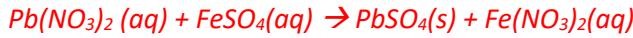
# Solutions

Soluble Ionic Compounds		Important Exceptions
Compounds containing	$\text{NO}_3^-$ $\text{C}_2\text{H}_5\text{O}_2^-$ $\text{Cl}^-$ $\text{Br}^-$ $\text{I}^-$ $\text{SO}_4^{2-}$	None None Compounds of $\text{Ag}^+$ , $\text{Hg}^{2+}$ , and $\text{Pb}^{2+}$ Compounds of $\text{Ag}^+$ , $\text{Hg}^{2+}$ , and $\text{Pb}^{2+}$ Compounds of $\text{Ag}^+$ , $\text{Hg}^{2+}$ , and $\text{Pb}^{2+}$ Compounds of $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Hg}^{2+}$ , and $\text{Pb}^{2+}$
Insoluble Ionic Compounds		Important Exceptions
Compounds containing	$\text{S}^{2-}$ $\text{CO}_3^{2-}$ $\text{PO}_4^{3-}$ $\text{OH}^-$	Compounds of $\text{NH}_4^+$ , the alkali metal cations, and $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , and $\text{Ba}^{2+}$ Compounds of $\text{NH}_4^+$ and the alkali metal cations Compounds of $\text{NH}_4^+$ and the alkali metal cations Compounds of the alkali metal cations, and $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , and $\text{Ba}^{2+}$

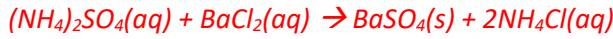
a) potassium hydroxide and lead(II) nitrate (Click for the solution to a)



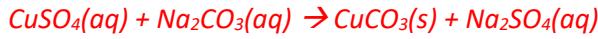
b) lead(II) nitrate and iron(II) sulfate



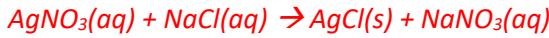
c) ammonium sulfate and barium chloride



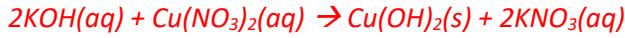
d) copper(II) sulfate and sodium carbonate



e) silver nitrate and sodium chloride



f) potassium hydroxide and copper(II) nitrate

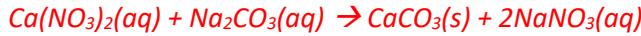


g) sodium hydroxide and iron(III) nitrate (Click for the solution to g)

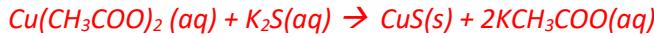
h) aluminium chloride and sodium hydroxide



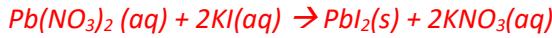
i) calcium nitrate and sodium carbonate



j) copper(II) acetate and potassium sulphide



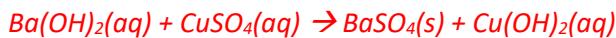
k) lead(II) nitrate and potassium iodide



l) magnesium sulfate and sodium hydroxide



m) Barium hydroxide and copper sulfate



n) Calcium sulfide and aluminium sulfate



*Two precipitates form.*