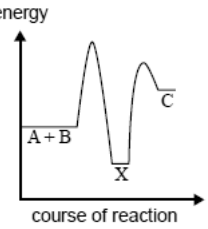
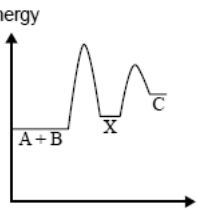
Thermo-chemistry 2007 VCE

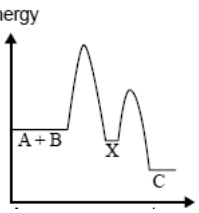
A chemical reaction has a ΔH of □-150 kJ mol$^{-□}$ and the activation energy for its reverse reaction is 350 kJ mol$^{-□}$.
The activation energy, in kJ mol$^{-□}$, of the forward reaction is
A. +500
B. +200
C. +150
D. -□200
Solution

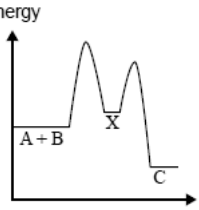
Solution will appear here

The reaction
A + B → C; ΔH negative
involves a two-step process
A + B → X; ΔH positive
X → C; ΔH negative
Which one of the following diagrams best represents the energy changes during the course of the reaction?

A.


B.


C.


D.


Solution will appear here

Solution

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