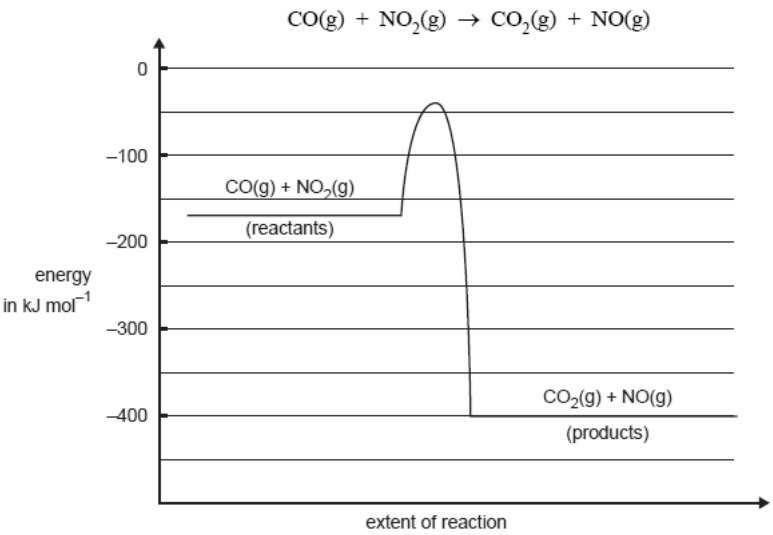
Thermochemistry (2005 VCE)

The graph below represents the energy changes over the course of a chemical reaction



Give the magnitude and sign of the ΔH for the forward reaction in kJ/mol□.
Solution

Give the activation energy for the reverse reaction in kJ mol□1.
Solution

Give two reasons explaining why the rate of this reaction increases with increasing temperature.
Solution

A suitable catalyst is discovered for the reaction. What would be the likely effect of the catalyst on:
- the activation energy? Explain your answer.
Solution

Solution will appear here

- the ΔH? Explain your answer.
Solution