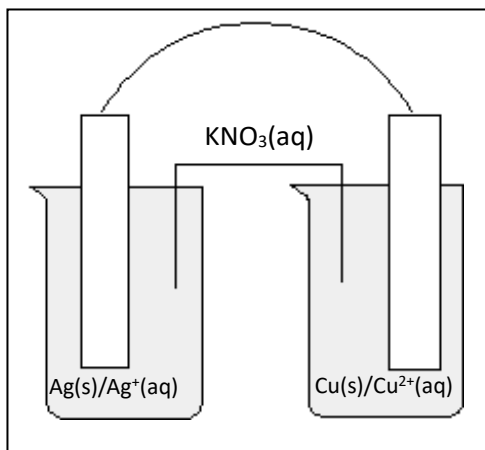


Video – worksheet – Labelling galvanic cells using the E° tables in the VCAA Data Booklet.

On the next page is a set of galvanic cells. Answer the questions by using the E° tables from the Data Booklet.

Refer to the video for solutions.

Use the blank space below to complete working out.



Show the following on the diagram.

Anode (-)

Cathode (+)

Electron flow.

Anion flow.

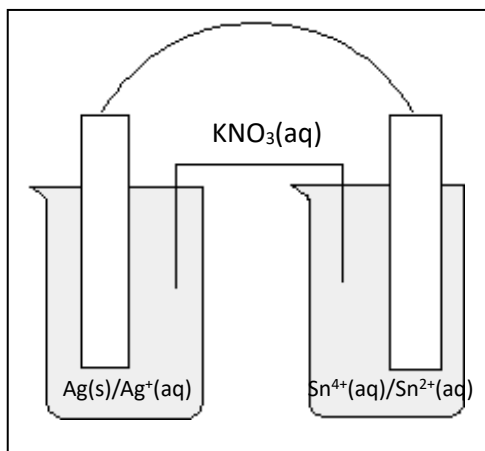
Oxidation half equation _____

Reduction half equation _____

Overall equation _____

What material is the anode made from? _____

What material is the cathode made from? _____



Show the following on the diagram.

Anode (-)

Cathode (+)

Electron flow.

Anion flow.

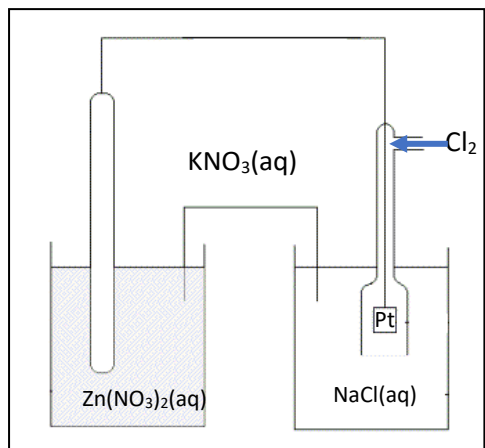
Oxidation half equation _____

Reduction half equation _____

Overall equation _____

What material is the anode made from? _____

What material is the cathode made from? _____



Show the following on the diagram.

Anode (-)

Cathode (+)

Electron flow.

Anion flow.

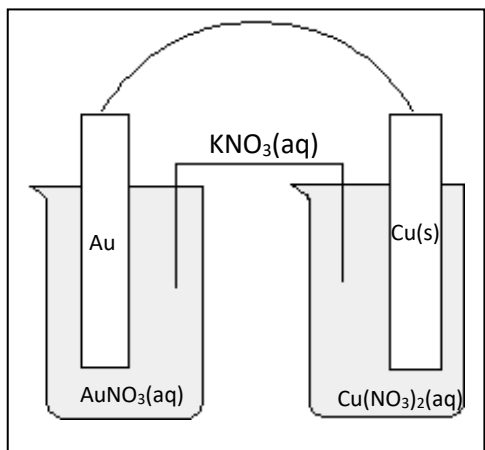
Oxidation half equation _____

Reduction half equation _____

Overall equation _____

What material is the anode made from? _____

What material is the cathode made from? _____



Show the following on the diagram.

Anode (-)

Cathode (+)

Electron flow.

Anion flow.

Oxidation half equation _____

Reduction half equation _____

Overall equation _____

What material is the anode made from? _____

What material is the cathode made from? _____