Acid reactions

- overall and ionic equations

Hydrogen	H^+	Chloride	Cl^-
Sodium	Na^+	Bromide	Br^-
Silver	Ag^+	Fluoride	F^-
Potasssium	K^+	Iodide	I^-
Lithium	Li^+	Hydroxide	OH^-
Ammonium	NH_4^+	Nitrate	NO_3^-
Barium	Ba^{2+}	Oxide	O^{2-}
Calcium	Ca^{2+}	Sulphide	S^{2-}
Copper(II)	Cu^{2+}	Sulphate	SO_4^{2-}
Magnesium	Mg^{2+}	Carbonate	CO_3^{2-}
Zinc	Zn^{2+}	Hydrogencarbonate	HCO_3^-
Lead	Pb^{2+}		
Iron(II)	Fe^{2+}		
Iron(III)	Fe^{3+}		
Aluminium	Al^{3+}		

Table 1
Valency of common ions

Reaction	
Sulphuric acid solution and aqueous sodium carbonate.	Overall : ionic :
Nitric acid solution and lithium metal	Overall : ionic :
Hydrochloric acid solution and magnesium sulphide powder	Overall : ionic :
Sulphuric acid solution and sodium oxide powder.	Overall : ionic :
Nitric acid solution and aqueous calcium hydroxide.	Overall : ionic :
Hydrochloric acid solution and aqueous sodium sulphite.	Overall : ionic :
Sulphuric acid solution and solid sodium hydroxide	Overall : ionic :
Nitric acid solution and aqueous sodium carbonate.	Overall : ionic :

Solutions